

AP Chemistry: A Sample Formal Laboratory Report

This paper is designed to help you prepare a chemistry lab report. Keep it in your chemistry notebook. All chemistry lab reports must be written neatly and well organized to receive full credit. Lab reports may be written or typed. It is highly recommended that you use graph or engineering bond paper for written reports.

Laboratory #7: Quantitative Determination of an Empirical Formula

I. Hypothesis: If nitric acid is poured onto tin, a tin oxide will be produced. If we know the initial mass of the tin metal and the mass of the final product, we can determine the empirical formula of the tin oxide product. There should be a whole number ratio between oxygen and tin. *(The hypothesis explains what is to be tested and will be written after reading the entire laboratory worksheet.)*

II. Equipment: *(Non-chemical equipment used in the experiment.)*
evaporating dish
watch glass
stirring rod
burner with ring stand, ring and wire gauze
forceps
beaker
balance

III. Reagents: *(A listing of chemicals used in the experiment with their amounts and any warnings.)*
tin metal (granulated) ~2 g.
5 cm³ (mL) nitric acid (HNO₃) **caution! severe burns**

IV. Procedure: Each step of the procedure must be written here. You may paraphrase and shorten the explanations, but the reader must be able to perform the experiment from these instructions. *(The procedure must be read carefully before the lab begins. Drawings of the experimental setup may be included here. The teacher may make changes to the procedure; make sure that you write any changes down!)*

V. Data: *(If the laboratory report is handwritten use a ruler to draw data tables and graphs! Always include units with all data entries.)*

	Procedure	Trial 1	Trial 2
a	mass of dish, and watch glass	74.14 g	
b	mass of dish, glass and tin	76.20 g	
c	mass of tin =b-a	2.06 g	
d	moles of tin	.0173 mol	
e	mass of dish, glass, and product	76.76 g	
f	mass of oxygen =e-b	.56 g	
g	moles of oxygen	.0350 mol	
h	mole ratio	2.02 : 1	
i	accepted ratio	2 : 1	
j	% error	1.00 %	

Examples Of Organic Chemistry Lab Reports

Robert Schaffer



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Organic Chemistry Lab I Experimental Procedures and Instructions Notebooks and Lab Reports - CHM 265, Fall 2013

Donald Boerth,2013 **Experimental Organic Chemistry** John C. Gilbert,Stephen F. Martin,2002-01-01

Determination of Organic Structures by Physical Methods E. A. Braude,F. C. Nachod,2013-10-22 Determination of Organic Structures by Physical Methods Volume 1 focuses on the processes methodologies principles and approaches involved in the determination of organic structures by physical methods including infrared light absorption thermodynamic properties Raman spectra and kinetics The selection first elaborates on the phase properties of small molecules equilibrium and dynamic properties of large molecules and optical rotation Discussions focus on simple acyclic compounds carbohydrates steroids diffusion viscosity osmotic pressure sedimentation velocity melting and boiling points and molar volume The book then examines ultraviolet and visible light absorption infrared light absorption Raman spectra and the theory of magnetic susceptibility Concerns cover applications to the study of organic compounds applications to the determination of structure determination of thermodynamic properties and experimental methods and evaluation of data The text ponders on wave mechanical theory reaction kinetics and dissociation constants including dissociation of molecular addition compounds principles of reaction kinetics and valence bond treatment of aromatic systems The selection is a valuable source of data for researchers interested in the determination of organic structures by physical methods **Contribution from the**

Research Laboratory of Organic Chemistry Massachusetts Institute of Technology. Research Laboratory of Organic Chemistry,1925 **Laboratory Experiments in Organic Chemistry** Ellen Parmelee Cook,1915 *Organic Chemistry Section: Summary of Activities July 1967 to June 1968* Robert Schaffer,1968 *Techniques in Organic Chemistry* Jerry R. Mohrig,Christina Noring Hammond,Paul F. Schatz,2010-01-06 Compatible with standard taper miniscale 14 10 standard

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Kremer,2018 Purification of Laboratory Chemicals W.L.F. Armarego,2003-03-07 Now in its fifth edition the book has been updated to include more detailed descriptions of new or more commonly used techniques since the last edition as well as remove those that are no longer used procedures which have been developed recently ionization constants pKa values and also more detail about the trivial names of compounds In addition to having two general chapters on purification procedures this book provides details of the physical properties and purification procedures taken from literature of a very extensive number of organic inorganic and biochemical compounds which are commercially available This is the only complete source that covers the purification of laboratory chemicals that are commercially available in this manner and format Complete update of this valuable well known reference Provides purification procedures of commercially available chemicals and

biochemicals Includes an extremely useful compilation of ionisation constants **Operational Organic Chemistry** John W. Lehman,1988 *Organic Chemistry Laboratory Operations* William Burns Renfrow,Peter Jack Hawkins,1962

Experiments in Organic Chemistry Louis Frederick Fieser,1935 *Organic Chemistry I Laboratory* Matt Davies,2015-11-11 **Write Like a Chemist** Marin Robinson,2008-08-18 Concise writing and organizational skills are stressed throughout and move structures teach students conventional ways to present their stories of scientific discovery

The Systematic Identification of Organic Compounds Ralph Lloyd Shriner,1980 Step by step instructions on identifying organic compounds The steps described include elemental analysis solubility infrared spectra nuclear magnetic resonance spectra mass spectra classification tests and preparation of a derivative Most directions for experiments are described in a micro or mini scale and clean up directions are given at the end of each procedure Emphasizes the systematic approach to identifying unknowns Offers a review of spectroscopy Discusses infrared nuclear magnetic resonance and mass spectroscopy and includes examples of spectra Discusses chromatography distillations and the separation of mixtures **Laboratory Technique in Organic Chemistry** Avery Adrian Morton,1973 **Organic Chemistry Section** United States. National Bureau of Standards,1969 **Report of Work Done in the Organic Chemical Laboratory** Frank H. Morgan,1878

Molecular Driving Forces Ken Dill,Sarina Bromberg,2010-10-21 *Molecular Driving Forces* Second Edition E book is an introductory statistical thermodynamics text that describes the principles and forces that drive chemical and biological processes It demonstrates how the complex behaviors of molecules can result from a few simple physical processes and how simple models provide surprisingly accurate insights into the workings of the molecular world Widely adopted in its First Edition *Molecular Driving Forces* is regarded by teachers and students as an accessible textbook that illuminates underlying principles and concepts The Second Edition includes two brand new chapters 1 Microscopic Dynamics introduces single molecule experiments and 2 Molecular Machines considers how nanoscale machines and engines work The Logic of Thermodynamics has been expanded to its own chapter and now covers heat work processes pathways and cycles New practical applications examples and end of chapter questions are integrated throughout the revised and updated text exploring topics in biology environmental and energy science and nanotechnology Written in a clear and reader friendly style the book provides an excellent introduction to the subject for novices while remaining a valuable resource for experts

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